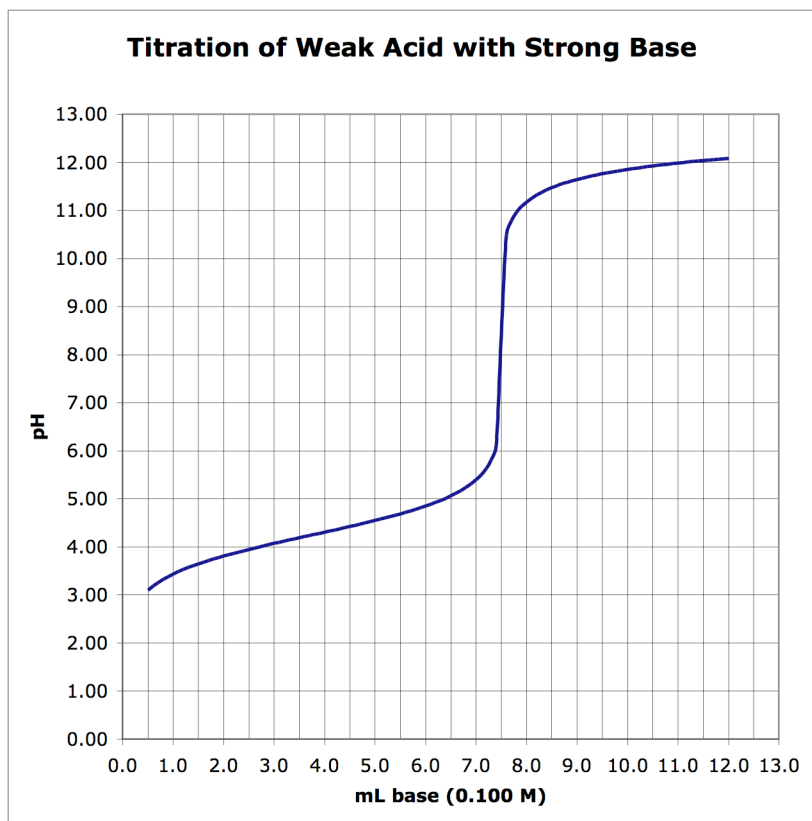


Titration of Weak Acids or Weak Bases



1) What is the pK_a of the weak acid?
 answer $pK_a = 4.3$

2) The equivalence point of this titration required how many mL of base?
 answer 7.5 mL

3) What was the original concentration of the acid if the starting volume of acid was 18.60 mL?
 answer $[HA] = 4.03 \times 10^{-2} \text{ M}$

4) What is the pK_b of the weak base?
 answer $pK_a = 10.5$ so $pK_b = 3.5$

5) The equivalence point of this titration required how many mL of acid?
 answer 10.0 mL

6) What was the original concentration of the base if the starting volume of base was 7.50 mL?
 $[B:] = 0.133 \text{ M}$

