

## Oxidation - Reduction Reactions

<u>Oxidations</u>	<u>Voltage</u>	<u>Reductions</u>	<u>Voltage</u>
$K \rightarrow K^+ + e^-$	+2.93	$F_2 + 2e^- \rightarrow 2F^-$	+2.87
$Ca \rightarrow Ca^{+2} + 2e^-$	+2.87	$2HClO + 2H^+ + 2e^- \rightarrow Cl_2 + 2H_2O$	+1.63
$Na \rightarrow Na^+ + e^-$	+2.71	$MnO_4^- + 8H^+ + 5e^- \rightarrow Mn^{+2} + 4H_2O$	+1.49
$Mg \rightarrow Mg^{+2} + 2e^-$	+2.36	$Cl_2 + 2e^- \rightarrow 2Cl^-$	+1.36
$Al \rightarrow Al^{+3} + 3e^-$	+1.66	$O_2 + 4H^+ + 4e^- \rightarrow 2H_2O$	+1.23
$Zn \rightarrow Zn^{+2} + 2e^-$	+0.76	$Br_2 + 2e^- \rightarrow 2Br^-$	+1.09
$Fe \rightarrow Fe^{+2} + 2e^-$	+0.44	$NO_3^- + 3H^+ + 2e^- \rightarrow HNO_2 + H_2O$	+0.94
$Ni \rightarrow Ni^{+2} + 2e^-$	+0.23	$I_2 + 2e^- \rightarrow 2I^-$	+0.54
$Sn \rightarrow Sn^{+2} + 2e^-$	+0.14	$O_2 + 2H_2O + 4e^- \rightarrow 4OH^-$	+0.40
$Fe \rightarrow Fe^{+3} + 3e^-$	+0.04	$SO_4^{-2} + 4H^+ + 2e^- \rightarrow H_2SO_3 + H_2O$	+0.20
$Cu \rightarrow Cu^{+2} + 2e^-$	-0.34	$2H^+ + 2e^- \rightarrow H_2$	0.00
$Ag \rightarrow Ag^+ + e^-$	-0.80	$S + 2e^- \rightarrow S^{-2}$	-0.48
$Hg \rightarrow Hg^{+2} + 2e^-$	-1.62	$2H_2O + 2e^- \rightarrow H_2 + 2OH^-$	-0.83