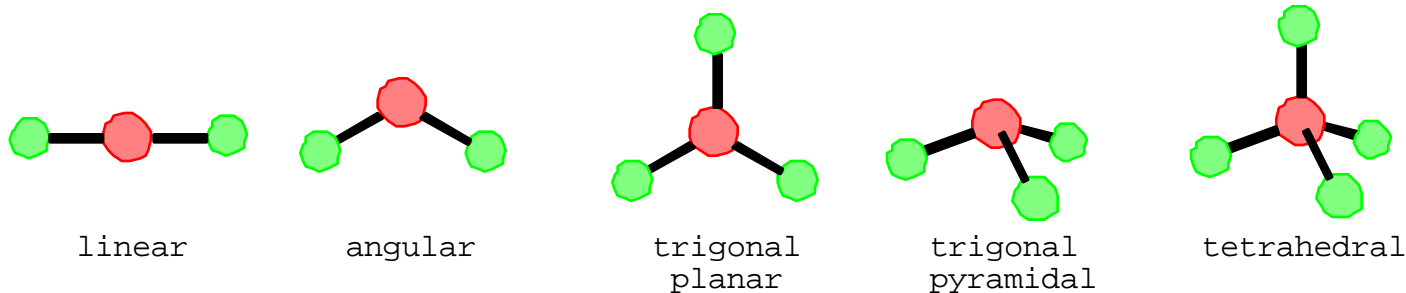


NAMES OF SIMPLE MOLECULAR SHAPES

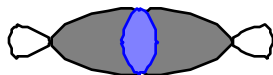
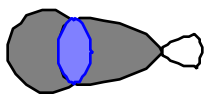
(these refer to atom arrangements around a central atom)



SIGMA AND PI BONDS

sigma bond (σ -bond)

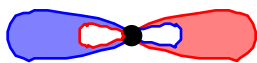
(overlap of s and hybrid orbital, or 2 hybrid orbitals along same axis)



pi bond (π -bond)

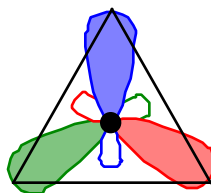
(overlap of parallel p orbitals)

HYBRIDIZATION OF ORBITALS



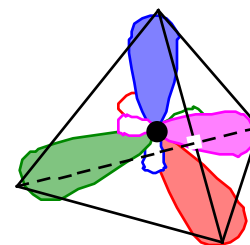
sp hybridization

(linear)



sp^2 hybridization

(trigonal planar)



sp^3 hybridization

(tetrahedral)