

Common Ions, Ionic Compounds

H^+									
Li^+	Be^{+2}						N^{-3}	O^{-2}	F^-
Na^+	Mg^{+2}					Al^{+3}		S^{-2}	Cl^-
K^+	Ca^{+2}								Br^-
Rb^+	Sr^{+2}					Ag^+			I^-
Cs^+	Ba^{+2}								

Mn^{+2}	Fe^{+2}	Co^{+2}	Ni^{+2}	Cu^+	Zn^{+2}
Mn^{+3}	Fe^{+3}	Co^{+3}		Cu^{+2}	
					Sn^{+2}
					Sn^{+4}
				Hg_2^{+2}	Pb^{+2}
Use the Stock System				Hg^{+2}	Pb^{+4}

SIMPLE IONS

POLYATOMIC IONS

NH_4^+	OH^-	$CH_3CO_2^-$	CO_3^{-2}	HCO_3^-
ammonium	hydroxide	acetate	carbonate	bicarbonate

NO_3^-	NO_2^-	PO_4^{-3}	HPO_4^{-2}	$H_2PO_4^-$
nitrate	nitrite	phosphate	monohydrogen phosphate	dihydrogen phosphate

SO_4^{-2}	SO_3^{-2}	OCl^-
sulfate	sulfite	hypochlorite

Total ion charges must be equal
for ionic compounds.

$$\# \oplus = \# \ominus$$

Hydrates - Ionic Compounds with attached water molecules in a definite proportion.

examples: $Na_2CO_3 \cdot 10H_2O$ $Co(NO_3)_2 \cdot 6H_2O$

PREFIXES

mono-	hexa-
di-	hepta-
tri-	octa-
tetra-	nona-
penta-	deca-