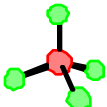

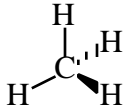
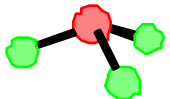
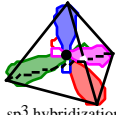
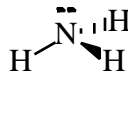
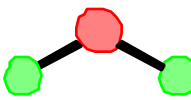

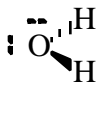
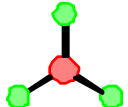

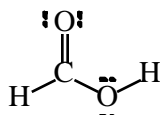
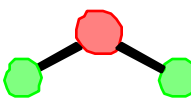
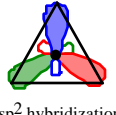
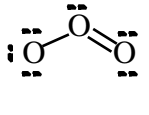

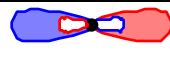
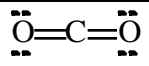


## Review of Molecular Shapes

Electron Areas	Bond Areas	Shape	Ball & Stick Model	Hybridization	Example
4	4	tetrahedral		 sp <sup>3</sup> hybridization	
4	3	trigonal pyramidal		 sp <sup>3</sup> hybridization	
4	2	angular		 sp <sup>3</sup> hybridization	
3	3	trigonal planar		 sp <sup>2</sup> hybridization	
3	2	angular		 sp <sup>2</sup> hybridization	
2	2	linear		 sp hybridization	

## Electronegativities

F	O	Cl	N	Br	I	C	S	H	P
4.0	3.4	3.2	3.0	3.0	2.7	2.6	2.6	2.2	2.2