

**Chem 105**  
**Practice Naming Compounds I**  
**Answers**

2) Ionic compounds - salts, acids and hydrates

LiCl lithium chloride

sodium sulfide  $\text{Na}_2\text{S}$

$\text{K}_2\text{O}$  potassium oxide

calcium bromide  $\text{CaBr}_2$

$\text{HBr(aq)}$  hydrobromic acid

sulfuric acid  $\text{H}_2\text{SO}_4(\text{aq})$

$\text{PbSO}_4$  lead(II)sulfate

iron(II)hydroxide  $\text{Fe(OH)}_2$

$\text{Mg}_3(\text{PO}_4)_2$  magnesium phosphate

potassium nitrite  $\text{KNO}_2$

$\text{NaC}_2\text{H}_3\text{O}_2 \cdot 3\text{H}_2\text{O}$  sodium acetate trihydrate

ammonium phosphate trihydrate  $(\text{NH}_4)_3\text{PO}_4 \cdot 3\text{H}_2\text{O}$

3) Binary Inorganic Molecular Compounds

$\text{N}_2\text{O}_5$  dinitrogen pentoxide

sulfur trioxide  $\text{SO}_3$

$\text{SiO}_2$  silicon dioxide

carbon disulfide  $\text{CS}_2$

$\text{PCl}_5$  phosphorus pentachloride

dichlorine monoxide  $\text{Cl}_2\text{O}$

4) Simple Organic Molecular Compounds

$\text{C}_3\text{H}_8$  propane

octane  $\text{C}_8\text{H}_{18}$

$\text{C}_5\text{H}_{11}\text{Cl}$  chloropentane

ethanol  $\text{C}_2\text{H}_5\text{OH}$

$\text{C}_9\text{H}_{19}\text{CO}_2\text{H}$  decanoic acid

hexene  $\text{C}_6\text{H}_{12}$